

Table 13.10 EXAMPLES OF TINCTURES APPLIED TO THE SKIN

TINCTURE	PERCENT ACTIVE CONSTITUENT IN COMMERCIAL TINCTURE	VEHICLE	CATEGORY AND COMMENTS
Green soap tincture	65%	Alcohol	Detergent. Also contains 2% lavender oil as perfume
Iodine tincture	2%	Alcohol, water	Topical anti-infective
Compound benzoin tincture	10% benzoin; 2% aloe; 8% storax; 4% Tolu balsam	Alcohol	Topical protectant. Prepared by maceration in alcohol
Podophyllin	Podocon-25	Benzoin tincture	Removal of soft genital warts

in these tables, most of these preparations are used as anti-infective agents. All medications intended for external use should be clearly labeled FOR EXTERNAL USE ONLY and kept out of the reach of children. In addition to their listing in Table 13.9, the following topical solutions are discussed because of their particular pharmaceutical interest.

Aluminum Acetate Topical Solution

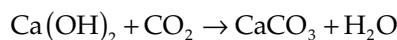
Aluminum acetate is colorless and has a faint acetous odor and a sweetish, astringent taste. It is widely applied topically as an astringent wash or wet dressing after dilution with 10 to 40 parts of water. It is frequently used in various types of dermatologic lotions, creams, and pastes. Commercial premeasured tablets and packets of powders are available for preparation of this solution. Synonym: Burow solution.

Aluminum Subacetate Topical Solution

The requirement for the amount of acetic acid differentiates aluminum acetate topical solution from aluminum subacetate topical solution. In the subacetate solution, the ratio of aluminum oxide to acetic acid is 1:2.35, whereas in the acetate solution, the ratio is 1:3.52. Aluminum subacetate topical solution, the stronger of the two, is used in preparation of aluminum acetate topical solution. Aluminum acetate topical solution, diluted first with 20 to 40 parts of water, is used externally as an astringent wash and wet dressing (modified Burow solution).

Calcium Hydroxide Topical Solution

Calcium hydroxide topical solution, commonly called limewater, must contain not less than 140 mg of $\text{Ca}(\text{OH})_2$ in each 100 mL of solution. Calcium hydroxide is less soluble in hot than in cold water, and cool purified water is the solvent. The solution is intended to be saturated with solute, and to ensure saturation, an excess of calcium hydroxide, 300 mg for each 100 mL of solution to be prepared, is agitated with the purified water, vigorously and repeatedly, for 1 hour. After this time, the excess calcium hydroxide is allowed to settle to the bottom of the container. This permits the solution to remain saturated should a portion of the dissolved solute at the solution's surface react with the carbon dioxide of the air to form insoluble calcium carbonate:



The calcium carbonate settles to the bottom of the container and, by appearance, is indistinguishable from the remaining excess of calcium hydroxide. The calcium hydroxide reserve dissolves as calcium is removed from the solution in the form of the carbonate, and, in this way, it continually maintains the saturation of the solution. After the solution stands for an appreciable length of time, the undissolved material at the bottom of the container is composed of varying proportions of calcium hydroxide and calcium carbonate. Because of the uncertainty of the residue's composition, one may not prepare additional quantities