

Grashof number $Gr = \frac{D^3 \rho^2 g \beta \Delta T}{\mu^2}$ (g : gravity acceleration, β : thermal expansion coefficient)

Prandtl number $Pr = \frac{c_p \mu}{k}$ (c_p : specific heat)

Graetz number (for hollow fiber membrane modules with length L and mass flowrate \dot{m}): $Gr = \frac{\dot{m} c_p}{kL}$

Since the membrane acts as hetero-nucleant interface with low surface, and due to concentration polarization that increases the solute concentration in proximity of the membrane, the nucleation process in a membrane crystallizer preferentially occurs on the membrane module; crystals are then flushed and recirculated into a separate crystallizer tank by the retentate flow. The combination of membrane modules and temperature-controlled crystallization should be considered as one system consisting of both nucleation and growth steps. Moreover, the retentate flowrate plays a crucial role in determining the detachment of crystals from the membrane surface and, in general, on the overall crystallization process.

For a continuous process, the growth rate (G) is estimated from the variation of the overall crystal length (\bar{L}) as a function of the residence time (τ):

$$G = \frac{d\bar{L}(\tau)}{d\tau} \approx \frac{\bar{L}(\tau_{n+1}) - \bar{L}(\tau_n)}{(\tau_{n+1} - \tau_n)} \quad (8.28)$$

where:

$$\bar{L}(\tau) \approx \sum_i \alpha_i(\tau) L_i \quad (8.29a)$$

$$\alpha_i(\tau) = \frac{N_i(\tau)}{N_{\text{total}}(\tau)} \quad (8.29b)$$

The crystal growth rate is empirically related to the supersaturation (S):

$$G = k_g S^g \quad (8.30)$$

where k_g is the growth rate constant (function of temperature, fluid-dynamics, presence of impurity *etc.*) and exponent g represents the kinetic order of the growth process. For NaCl crystallization on microporous polypropylene membranes carried out at low supersaturation ratio ($S = 0.065-0.097$), least square multiple linear regression analysis of data resulted in $g = 1.17$ and $k_g = 1.92 \times 10^{-5}$ (G expressed in m s^{-1}).²