

obtained. Evaporate about 15 mL of the filtrate to about 1 mL on a hot plate. Deposit this solution on a silver chloride disk.

**Standard specimen:** a similar preparation of polyethylene glycol.

**pH** (791): between 3.0 and 5.0, in a solution [15 in 100 (w/w)].

**Arsenic, Method I** (211): 2 µg per g.

**Delete the following:**

**Heavy metals, Method I** (231): 20 µg per g. (Official 1-Jan-2018)

**Limit of iron**—Using Aluminum Dichlorohydrate Polyethylene Glycol instead of Aluminum Chlorohydrate, proceed as directed in the test for *Limit of iron* under *Aluminum Chlorohydrate*. The limit is 150 µg per g.

**Content of aluminum**—Using Aluminum Dichlorohydrate Polyethylene Glycol instead of Aluminum Chlorohydrate, proceed as directed in the test for *Content of aluminum* under *Aluminum Chlorohydrate*. Use the result obtained to calculate the *Aluminum/chloride atomic ratio*.

**Content of chloride**—Using Aluminum Dichlorohydrate Polyethylene Glycol instead of Aluminum Chlorohydrate, proceed as directed in the test for *Content of chloride* under *Aluminum Chlorohydrate*. Use the result obtained to calculate the *Aluminum/chloride atomic ratio*.

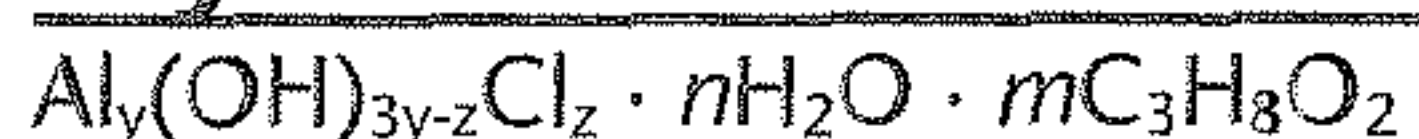
**Aluminum/chloride atomic ratio**—Divide the percentage of aluminum found in the test for *Content of aluminum* by the percentage of chloride found in the test for *Content of chloride*, and multiply by 35.453/26.98, in which 35.453 and 26.98 are the atomic weights of chlorine and aluminum, respectively: the ratio is between 0.90:1 and 1.25:1.

**Assay**—Calculate the percentage of anhydrous aluminum dichlorohydrate in the Aluminum Dichlorohydrate Polyethylene Glycol by the formula:

$$Al\{26.98x + [17.01(3x - 1)] + 35.453\} / 26.98x$$

in which *Al* is the percentage of aluminum found in the test for *Content of aluminum*, *x* is the aluminum-to-chloride atomic ratio, 26.98 is the atomic weight of aluminum, 17.01 is the molecular weight of the hydroxide anion (OH), and 35.453 is the atomic weight of chlorine (Cl).

## Aluminum Dichlorohydrate Propylene Glycol



Aluminum chlorohydroxide propylene glycol complex.

Aluminum hydroxychloride propylene glycol complex.

» Aluminum Dichlorohydrate Propylene Glycol consists of aluminum dichlorohydrate in which some of the waters of hydration have been replaced by propylene glycol. It encompasses a range of aluminum-to-chloride atomic ratios between 0.90:1 and 1.25:1. It contains not less than 90.0 percent and not more than 110.0 percent of the labeled amount of anhydrous aluminum dichlorohydrate.

**Packaging and storage**—Preserve in well-closed containers.

**Labeling**—The label states the content of anhydrous aluminum dichlorohydrate.

## Identification—

**A:** A solution (1 in 10) responds to the tests for *Aluminum* (191) and for *Chloride* (191).

**B:** Dissolve 0.5 g in about 40 mL of water, and while mixing adjust with 2.5 N sodium hydroxide to a pH of 9.55 ± 0.05. Filter the suspension of precipitate thus obtained. Evaporate about 15 mL of the filtrate to about 1 mL on a hot plate: the IR absorption spectrum of a film of this solution on a silver chloride disk exhibits maxima only at the same wavelengths as that of a similar preparation of a film of propylene glycol.

**pH** (791): between 3.0 and 5.0, in a solution [15 in 100 (w/w)].

**Arsenic, Method I** (211): 2 µg per g.

**Delete the following:**

**Heavy metals, Method I** (231): 20 µg per g. (Official 1-Jan-2018)

**Limit of iron**—Using Aluminum Dichlorohydrate Propylene Glycol instead of Aluminum Chlorohydrate, proceed as directed in the test for *Limit of iron* under *Aluminum Chlorohydrate*. The limit is 150 µg per g.

**Content of aluminum**—Using Aluminum Dichlorohydrate Propylene Glycol instead of Aluminum Chlorohydrate, proceed as directed in the test for *Content of aluminum* under *Aluminum Chlorohydrate*. Use the result obtained to calculate the *Aluminum/chloride atomic ratio*.

**Content of chloride**—Using Aluminum Dichlorohydrate Propylene Glycol instead of Aluminum Chlorohydrate, proceed as directed in the test for *Content of chloride* under *Aluminum Chlorohydrate*. Use the result obtained to calculate the *Aluminum/chloride atomic ratio*.

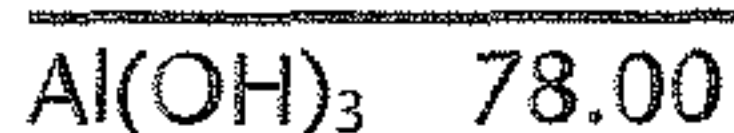
**Aluminum/chloride atomic ratio**—Divide the percentage of aluminum found in the test for *Content of aluminum* by the percentage of chloride found in the test for *Content of chloride*, and multiply by 35.453/26.98, in which 35.453 and 26.98 are the atomic weights of chlorine and aluminum, respectively: the ratio is between 0.90:1 and 1.25:1.

**Assay**—Calculate the percentage of anhydrous aluminum dichlorohydrate in the Aluminum Dichlorohydrate Propylene Glycol by the formula:

$$Al\{26.98x + [17.01(3x - 1)] + 35.453\} / 26.98x$$

in which *Al* is the percentage of aluminum found in the test for *Content of aluminum*, *x* is the aluminum-to-chloride atomic ratio, 26.98 is the atomic weight of aluminum, 17.01 is the molecular weight of the hydroxide anion (OH), and 35.453 is the atomic weight of chlorine (Cl).

## Aluminum Hydroxide Gel



Aluminum hydroxide.

Aluminum hydroxide [21645-51-2].

» Aluminum Hydroxide Gel is a suspension of amorphous aluminum hydroxide in which there is a partial substitution of carbonate for hydroxide. It contains the equivalent of not less than 90.0 percent and not more than 110.0 percent of the labeled amount of aluminum hydroxide [Al(OH)<sub>3</sub>]. It may contain Peppermint Oil, Glycerin, Sorbitol, Sucrose, Saccharin, or other suitable flavors, and it may contain suitable antimicrobial agents.